|  |
| --- |
| **Faculty of Sciences and Mathematics, UNIVERSITY OF NIŠ** |
| **Course Unit Descriptor** | **Faculty** |  |
| **GENERAL INFORMATION** |
| Study program  | **Undergraduate** |
| Study Module (if applicable) |  |
| Course title | General Chemistry |
| Level of study | ☒ Bachelor ☐ Master’s ☐ Doctoral |
| Type of course | ☒ Obligatory☐ Elective |
| Semester  | ☒ Autumn ☐Spring |
| Year of study  | First |
| Number of ECTS allocated | 10 |
| Name of lecturer/lecturers | Nikola D. Nikolić, Maja N. Stanković |
| Teaching mode | ☒ Lectures ☐Group tutorials ☐ Individual tutorials☒ Laboratory work ☐ Project work ☐ Seminar☐ Distance learning ☐ Blended learning ☐ Other |
| **PURPOSE AND OVERVIEW (max. 5 sentences)** |
| *Introduce students to basic materials in chemistry and adoption of general chemical laws in order of successful attempting of subsequent modules**Acquisition of the knowledge about the structure of atoms, chemical bonding, reactions, types of inorganic compounds and their systematization.**Acquisition of basic knowledge of nomenclature, chemical bonds, solution behaviour, thermochemistry, kinetics and electrochemistry* |
| **SYLLABUS (brief outline and summary of topics, max. 10 sentences)** |
| **Atomic orbitals. Quantum numbers and levels. Magnetic properties of atoms and ions. Physicochemical properties of elements (electronic affinity, ionisation potentials, ionic radius). The theory of chemical bonding. Geometry of molecules, dipole moment. Solutions. Oxidation-reduction reactions. Chemical kinetic and equilibrium. Complex compounds.** |
| **LANGUAGE OF INSTRUCTION** |
| ☒ Serbian (complete course) ☐ English (complete course) ☐ Other \_\_\_\_\_\_\_\_\_\_\_\_\_ (complete course)☐Serbian with English mentoring ☐Serbian with other mentoring \_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| **ASSESSMENT METHODS AND CRITERIA** |
| **Pre exam duties** | **Points** | **Final exam** | **points** |
| **Activity during lectures** | **5** | **Written examination** | **20** |
| **Practical teaching** | **5** | **Oral examination** | **20** |
| **Teaching colloquia** | **50** | **OVERALL SUM** | **100** |
| **\*Final examination mark is formed in accordance with the Institutional documents** |